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$\ln K = \left(\frac{\Delta H^\circ}{R} \right) - \left(\frac{\Delta H^\circ}{RT} \right)$ The equilibrium for the dissociation of one mole of iodine vapour is 3.94×10^{-4} at 900K and the standard enthalpy is 154 kJmol⁻¹. Calculate the equilibrium constant at 1200K assuming the standard enthalpy remains constant. $\ln (K_2/K_1) = - \frac{\Delta H^\circ}{R} \left(\frac{1}{T_2} - \frac{1}{T_1} \right)$, ANS = 6.7×10^{-2} at 1200K.

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