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Solution: Let a be the length of the sides of the tetrahedral cube and r be the radius of the two types of atoms.

$$a^3 = r^3 + r^3 \quad C \quad H \quad a = 2r_C + 2r_H = 2$$
$$(0.77 \times 10^{-8} \text{ cm} + 0.46 \times 10^{-8} \text{ cm}) = 1.42 \times 10^{-8} \text{ cm}$$

7.2–20 The compound aluminum phosphide (AlP) is a compound semiconductor having mixed ionic and covalent bonding.

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