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$l H P_{vap} = \exp$

$14.790 \cdot T / K$  and by

the ideal gas law

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$N V R T \times 18 \text{ mol} \times$

$8.314 \times 10^3 \text{ (bar} \cdot \text{m}^3 \text{ mol}^{-1} \text{K}^{-1})$

$T / K$  )  $T / P = =$

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$VV_1 \times 10^3 \text{ ? } (10 \text{ ? } x)$

106 volume taken up

by liquid Equating P

and P vap we have x

$18 \times 8.314 \times 10^5 T F$

$5432.8 \times l .$

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